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Sorption properties of composite materials with a core-shell structure containing layered double hydroxides in the shell

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Abstract

The work is devoted to the study of the sorption properties of hierarchical composite materials with a core-shell structure. The composites contained a core of SiO₂ or Fe₃O₄(*a*)SiO₂ obtained by sol-gel synthesis, on the surface of which a layered double hydroxide (MgAlFe-LDH) was deposited. The phase composition of the obtained materials was determined, and the textural characteristics and particle morphology were studied. It was found that hierarchical materials had larger surface and demonstrated high sorption capacity towards both cationic and anionic dyes in aqueous solution in comparison with individual systems (SiO₂ and MgAlFe-LDH). It was shown that the sorption equilibrium in the system "dye solution - sorbent" for dye methylene blue was achieved faster in comparison with Congo red. The obtained kinetic data were analyzed using chemical kinetic models. The sorption of both Congo red and methylene blue on composite materials was found to be described by a pseudo-second order kinetic equation. Isotherms of sorption of Congo red and methylene blue on synthesized materials were plotted. The sorption capacity of Fe₃O₄@SiO₂@LDH and SiO2@LDH towards Congo red were 0.19 mmol/g and 0.27 mmol/g, respectively. In the case of sorption of methylene blue, the sorption isotherms did not reach a plateau in the studied concentration range. However, it can be noted that at an initial methylene blue concentration of 0.051 mmol/L the sorption capacity of Fe₃O₄@SiO₂@LDH and SiO₂@LDH were 0.040 mmol/g and 0.033 mmol/g, respectively. The obtained data indicate that hierarchical composite materials containing LDH in their composition are effective bifunctional sorbents and can uptake both anionic and cationic forms of pollutants from a solution. An advantage of the Fe₃O₄ core system is its ability to be easily separate from a solution under the influence of an external magnetic field. It is important that the Fe₃O₄@SiO₂@LDH sample exhibits a typical superparamagnetic behavior with zero coercitivity and residual magnetic induction.

References

- [1] R. Sivashankar, A.B. Sathya, K. Vasantharaj, V. Sivasubramanian. Magnetic composite an environmental super adsorbent for dye sequestration – A review. Environmental Nanotechnology, Monitoring & Management. 2014. Vol.1-2. P.36-49.
- [2] M. Ruthiraan, M.N. Mubarak, E.C. Abdullah, M. Khalid, S. Nizamuddin, R. Walvekar, R.R. Karri. An Overview of Magnetic Material: Preparation and Adsorption Removal of Heavy Metals from Wastewater. In: Abd-Elsalam K., Mohamed M., Prasad R. (Eds.) Magnetic Nanostructures. Nanotechnology in the Life Sciences. Springer, Cham. 2019. P.131-159.
- [3] O.V. Makarchuk, T.A. Dontsova, I.M. Astrelin. Magnetic Nanocomposites as Efficient Sorption Materials for Removing Dyes from Aqueous Solutions. Nanoscale Research Letters. 2016. Vol.11. P.161.
- [4] P.T. Minh, O.E. Lebedeva. Adsorption Properties of a Magnetite Composite with Coffee Waste. Russian Journal of Physical Chemistry A. 2018. Vol.92. Iss.10. P.2044-2047.
- [5] R. Shan, L. Yan, Y. Yang, K. Yang, S. Yu, H. Yu, B. Zhu, B. Du. Highly efficient removal of three red dyes by adsorption onto Mg-Al-layered double hydroxide. Journal of Industrial and Engineering Chemistry. 2015. Vol.21. P.561-568.

SORPTION PROPERTIES OF COMPOSITE MATERIALS WITH A CORE-SHELL STRUCTURE CONTAINING... 62-71

- [6] R. Lafi, K. Charradi, M.A. Djebbi, A. Ben Haj Amara, A. Hafiane. Adsorption study of Congo red dye from aqueous solution to Mg-Al-layered double hydroxide. Advanced Powder Technology. 2016. Vol.27. Iss.1. P.232-237.
- [7] I. Ryltsova, E. Tarasenko, O. Nestroinaia, O. Lebedeva Sorption properties of MgCo/AlFe layered double hydroxides. Sorption and Chromatographic Processes. 2019. Vol.19. Iss.3. P.305-314. (russian)
- [8] K.-H. Goh, T.-T. Lim, Z. Dong. Application of layered double hydroxides for removal of oxyanions: A review. Water Research. 2008. Vol.42. P.1343-1368.
- [9] X. Liang, Y. Zang, Y. Xu, X. Tan, W. Hou, L. Wang, Y. Sun. Sorption of metal cations on layered double hydroxides. Colloids and Surfaces A: Physicochemical and Engineering Aspects. 2013. Vol.433. P.122-131.
- [10] I.G. Ryltsova, V.A. Piskareva, O.A. Vorontsova, and O.E. Lebedeva. Study of kinetics of Fe²⁺ sorption from aqueous solution by layered double hydroxides with hydrotalcite structure. Butlerov Communications. 2016. Vol.48. No.10. P.77-82. https://doi.org/10.37952/ROI-jbc-02/16-48-10-77
- [11] F. Cavani, F. Trifirò, A. Vaccari. Hydrotalcite-type anionic clays: preparation, properties and applications. Catal. Today. 1991. Vol.11. P.173-301.
- [12] D.G. Evans, R.C.T. Slade. Structural aspects of layered double hydroxides. In: X. Duan, D.G. Evans (Eds.), Layered Double Hydroxides. Structure and Bonding. Springer, Berlin, Heidelberg. 2006. Vol.119. P.1-87.
- [13] R. Shan, L. Yan, K. Yang, S. Yu, Y. Hao, H. Yu, B. Du. Magnetic Fe₃O₄/MgAl-LDH composite for effective removal of three red dyes from aqueous solution. Chemical Engineering Journal. 2014. Vol.252. P.38-46.
- [14] D. Chen, Y. Li, J. Zhang, J. Zhou, Y. Guo, H. Liu. Magnetic Fe₃O₄/ZnCr-layered double hydroxide composite with enhanced adsorption and photocatalytic activity. *Chemical Engineering Journal.* 2012. Vol.185-186. P.120-126.
- [15] P. Koilraj, K. Sasaki. Fe₃O₄/MgAl-NO₃ layered double hydroxide as a magnetically separable sorbent for the remediation of aqueous phosphate. Journal of Environmental Chemical Engineering. 2016. Vol.4. Iss.1. P.984-991.
- [16] R. Wang, T. Wen, X. Wu, A. Xu. Highly efficient removal of humic acid from aqueous solutions by Mg/Al layered double hydroxides-Fe3O4 nanocomposites. RSC Adv. 2014. Vol.4. Iss.42. P.21802-21809.
- [17] H. Palza, K. Delgado, J. Govan. Novel magnetic CoFe₂O₄/layered double hydroxide nanocomposites for recoverable anionic adsorbents for water treatment. Applied Clay Science. 2019. Vol.183. Article: 105350.
- [18] Q. Yan, Z. Zhang, Y. Zhang, A. Umar, Z. Guo, D. O'Hare, Q. Wang. Hierarchical Fe₃O₄Core–Shell Layered Double Hydroxide Composites as Magnetic Adsorbents for Anionic Dye Removal from Wastewater. European Journal of Inorganic Chemistry. 2015. P.4182-4191.
- [19] L. Lu, J. Li, D.H.L. Ng, P. Yang, P. Song, M. Zuo. Synthesis of novel hierarchically porous Fe₃O₄@MgAl-LDH magnetic microspheres and its superb adsorption properties of dye from water. Journal of Industrial and Engineering Chemistry. 2017. Vol.46. P.315-323.
- [20] W. Stober, A. Fink, E. Bohn. Controlled Growth of Monodisperse Silica Spheres in the Micron Size Range. Journal of Colloid and Interface Science. 1968. Vol.26. P.62-69.
- E.A. Tarasenko, I.G. Ryl'tsova, M.N. Yapryntsev, D.E. Smalchenko, O.E. Lebedeva. Synthesis of a [21] Magnetic Core/Shell Nanocomposite Containing Layered Double Hydroxide. Petroleum Chemistry. 2019. Vol.59. Iss.8. P.875-879.
- [22] Грег С., Синг К. Адсорбция, удельная поверхность, пористость. Пер. с англ. 2-е изд. Москва, Mup. 1984. 306c. [S.J. Gregg, K.S.W. Sing Adsorption, surface area and porosity. Translation from English. Second Edition. Moscow, Mir. 1984. 306 p. (russian)]
- [23] Y. Hou, J. Yu, S. Gao. Solvothermal reduction synthesis and characterization of superparamagnetic magnetite nanoparticles. Journal of Materials Chemistry. 2003. Vol.13. P.1983-1987.